

Endothermic and Exothermic Reactions

In-During a chemical reaction, the making and breaking of chemical bonds converts ~~the~~ reactants into products. When two or more ~~such substances~~ elements ~~counter~~ react, a chemical bond ~~is~~ formed between/among atoms and forms a chemical compound. Chemical reactions are classified into two types: endothermic and exothermic reactions. Energy activation results in the bonding of ~~the~~ two reactants to form a new product. All chemical reactions are accompanied by a change in energy.

Comment [A1]: "During" indicates a period or range of time (having duration) and is used to say that something happened. A clause with during focuses more on what happened—the activity, event, or experience. A clause with "in" focuses more on when something happened rather than what happened.

~~There are m~~ Many chemical reactions release emit energy with-in the forms of heat, light, or sound. Such ~~chemical~~ reactions are called exothermic reactions. ~~This~~ The released energy that is released comes originates from the bonds that join-link several atoms together in the molecules. A common example of exothermic reactions is ~~the phenomena of~~ combustion. ~~A fully combustible~~ Complete combustion process is occurs when a compound reacts with an oxidizing substance/element, yielding and the compounds of each element in the fuel with the oxidizing element are emitted as products. ~~There~~ Most exothermic reactions are mostly spontaneous exothermal processes.

~~On the other hand~~ Conversely, many other chemical reactions absorb energy in the form of heat, light, or sound. Such ~~chemical~~ reactions are called endothermic reactions. ~~These~~ reactions cannot progress with proceed without addition of heat or supplying energy. The resulting product of ~~the~~ such a reaction is less stable because, a molecule becomes less stable as the energy of its bonds increases the higher the energy bond, the less strength its molecules possess. A common example of endothermic reactions is ~~the~~ phenomena of photosynthesis. Here, plants use ~~the~~ energy from the sun to convert carbon dioxide and water into glucose and oxygen. Most endothermic reactions are not spontaneous.

Comment [A2]: "On the other hand" is used after the phrase "On the one hand" when comparing two different facts or two opposite ways. "Conversely" is a better word choice at this instance as "On the one hand" has not been used.

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To understand the difference between the two types of reactions, we need to explore several concepts

~~like such as the behavior of~~ the kinetic energy and potential energy ~~behavior inof~~ the reactant molecules
~~of the reactants of the chemical reaction.~~

Comment [A3]: The original sentence was unclear and needed complete rewriting to make the sentence unambiguous. Redundancies (“of the reactants of the chemical reaction”) have been removed and appropriate words (“to understand” instead of “to know”) have been used.

SAMPLE